

Ionic Compound Nomenclature Worksheet #3

Write the formula for each compound. Be SURE to include charges on each ion.				
1.	ammonium phosphate	$(\text{NH}_4^+)_3\text{PO}_4^{3-}$	16. beryllium borate	$\text{Be}_3(\text{BO}_3^3)^2$
2.	silver chloride	Ag^+Cl^-	17. strontium phosphide	$\text{Sr}_3^{2+}\text{P}_2^{3-}$
3.	barium nitrate	$\text{Ba}^{2+}(\text{NO}_3^-)_2$	18. barium manganate	$\text{Ba}^{2+}\text{MnO}_4^{2-}$
4.	aluminum sulfite	$\text{Al}_2^{3+}(\text{SO}_3^{2-})_3$	19. lead(II) amide	$\text{Pb}^{2+}(\text{NH}_2^-)_2$
5.	rubidium permanganate	$\text{Rb}^+\text{MnO}_4^-$	20. ammonium bromide	NH_4^+Br^-
6.	lithium phosphite	$\text{Li}_3^+\text{PO}_3^{3-}$	21. calcium ferrocyanide	$\text{Ca}_2^{2+}\text{Fe}(\text{CN})_6^{4-}$
7.	calcium fluoride	$\text{Ca}^{2+}\text{F}_2^-$	22. copper(II) phosphate	$\text{Cu}_3^{2+}(\text{PO}_4^{3-})_2$
8.	iron(III) nitrate	$\text{Fe}^{3+}(\text{NO}_3^-)_3$	23. aluminum bromate	$\text{Al}^{3+}(\text{BrO}_3^-)_3$
9.	zinc acetate	$\text{Zn}^{2+}(\text{C}_2\text{H}_3\text{O}_2^-)_2$	24. lithium chloride	Li^+Cl^-
10.	cobalt(II) hydroxide	$\text{Co}^{2+}(\text{OH}^-)_2$	25. potassium sulfite	$\text{K}_2^+\text{SO}_3^{2-}$
11.	magnesium nitride	$\text{Mg}_3^{2+}\text{N}_2^{3-}$	26. nickel(II) azide	$\text{Ni}^{2+}(\text{N}_3^-)_2$
12.	chromium(IV) acetate	$\text{Cr}^{4+}(\text{C}_2\text{H}_3\text{O}_2^-)_4$	27. silver phosphide	$\text{Ag}_3^+\text{P}^{3-}$
13.	aluminum oxide	$\text{Al}_2^{3+}\text{O}_3^{2-}$	28. strontium aluminate	$\text{Sr}_3^{2+}(\text{AlO}_3^3)^2$
14.	calcium thiocyanate	$\text{Ca}^{2+}(\text{SCN}^-)_2$	29. lithium fluoride	Li^+F^-
15.	cesium tartrate	$\text{Cs}_2^+\text{C}_4\text{H}_4\text{O}_6^{2-}$	30. aluminum bromide	$\text{Al}^{3+}\text{Br}_3^-$