

Quiz #27: VSEPR

Name: _____

Chemistry – Block _____

Date: _____

For each of the following, do the VSEPR arithmetic, draw the electron pair arrangement, then attach the atoms and name the molecular shape.

<p>CS₂ How many e⁻ pairs? <u> 2 </u></p> <p>Drawing:</p> <p>Name of molecular shape: LINEAR</p>	<p>PH₃ How many e⁻ pairs? <u> 4 </u></p> <p>Drawing:</p> <p>Name of molecular shape: shape: TRIGONAL PYRAMID</p>	<p>H₂O How many e⁻ pairs? <u> 4 </u></p> <p>Drawing:</p> <p>Name of molecular shape: ANGULAR</p>
<p>CCl₄ How many e⁻ pairs? <u> 4 </u></p> <p>Drawing:</p> <p>Name of molecular shape: TETRAHEDRON</p>	<p>SO₂⁻² How many e⁻ pairs? <u> 4 </u></p> <p>Drawing:</p> <p>Name of molecular shape: shape: ANGULAR</p>	<p>BCl₃ How many e⁻ pairs? <u> 3 </u></p> <p>Drawing:</p> <p>Name of molecular shape: TRIGONAL PLANAR</p>