

QUIZ #17

Pages 133-141 Reading

Directions: Choose the letter of the BEST answer and mark it onto your scantron sheet.

1. You can think of energy levels and sublevels much the same way you can visualize a section of seats in a theater. There is only 1 sublevel in level 1. There are how many sublevels in level 2?
 - A. None
 - B. One
 - C. Two
 - D. Three
 - E. Unknown
2. There are how many different **p** orbitals?
 - A. None
 - B. One
 - C. Two
 - D. Three
 - E. Unknown
3. Which of the following is NOT one of the orbital types?
 - A. s
 - B. p
 - C. a
 - D. d
 - E. f
4. The arrangement of electrons in an atom is called the atom's:
 - A. Energy level
 - B. Electron configuration
 - C. Aufbau
 - D. Electron sublevel
 - E. Electron energy order
5. The _____ states that a maximum of two electrons may occupy a single atomic orbital.
 - A. Hund's rule
 - B. Heisenburg uncertainty principle
 - C. Aufbau principle
 - D. Freije principal
 - E. Pauli exclusion principle
6. _____ states that single electrons with the same spin must occupy a different orbital before a second electron enters.
 - A. Hund's rule
 - B. Heisenburg uncertainty principle
 - C. Aufbau principle
 - D. Freije principal
 - E. Pauli exclusion principle

7. What is the special name that we give to electrons which are found in the outermost energy level or outer shell?
- A. Electron dot electrons
 - B. Aufbau electrons
 - C. Valence electrons
 - D. Orbital electrons
 - E. Amazing electrons
8. The electron dot structure shows the symbol of the atom with its:
- A. Electron dot electrons
 - B. Aufbau electrons
 - C. Valence electrons
 - D. Orbital electrons
 - E. Amazing electrons