

## ***Precipitation Reactions Table***

For each **RED NUMBER** below, write

**PPT**: if a precipitate **DOES** form

**NO PPT**: if **NO** precipitate forms

**PARTIAL PPT**: if “some” precipitate forms

<b>Compound</b>	$(\text{NH}_4)_2\text{SO}_4$	$\text{CaCl}_2$	$\text{NaOH}$	$\text{Pb}(\text{NO}_3)_2$	$\text{AgNO}_3$	$\text{K}_2\text{CO}_3$	$\text{Cu}(\text{C}_2\text{H}_3\text{O}_2)_2$	$\text{HCl}$
$(\text{NH}_4)_2\text{SO}_4$								
$\text{CaCl}_2$	<b>1</b>							
$\text{NaOH}$	<b>2</b>	<b>3</b>						
$\text{Pb}(\text{NO}_3)_2$	<b>4</b>	<b>5</b>	<b>6</b>					
$\text{AgNO}_3$	<b>7</b>	<b>8</b>	<b>9</b>	<b>10</b>				
$\text{K}_2\text{CO}_3$	<b>11</b>	<b>12</b>	<b>13</b>	<b>14</b>	<b>15</b>			
$\text{Cu}(\text{C}_2\text{H}_3\text{O}_2)_2$	<b>16</b>	<b>17</b>	<b>18</b>	<b>19</b>	<b>20</b>	<b>21</b>		
$\text{HCl}$	<b>22</b>	<b>23</b>	<b>24</b>	<b>25</b>	<b>26</b>	<b>27</b>	<b>28</b>	