

Density LAB

Name: _____

Date: _____

Block: _____

Procedure	Record below. Be sure to include correct units for EACH measurement!
1. Mass of 10 mL EMPTY graduated cylinder	
2. Mass of 10 mL graduated cylinder PLUS 10 mL dH ₂ O	
3. Calculate the mass of the 10 mL dH ₂ O by itself (without the graduated cylinder)	
4. Calculate the density of the distilled water ($D = m / V$) – be sure to use correct units!	
1. Mass of 250 mL EMPTY beaker	
2. Add 5.00 g of NaCl to beaker – how do you do this?	
3. Using graduated cylinder, measure out exactly 50 mL of dH ₂ O.	
4. Completely dissolve the 5.00 g NaCl into the 50 mL of dH ₂ O in the beaker.	
5. Mass of 10 mL EMPTY graduated cylinder	
6. Fill graduated cylinder with exactly 10 mL of the salt water.	
7. Measure mass of 10 mL salt water plus the graduated cylinder.	
8. Mass of the salt water by itself.	
9. Calculate the density of the salt water – be sure to use correct units!	
10. Which is more dense? dH ₂ O or salt water?	